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<b>(21) International Application Number:</b> PCT/EP95/01930 <b>(22) International Filing Date:</b> 19 May 1995 (19.05.95) <b>(30) Priority Data:</b> 94201680.9 13 June 1994 (13.06.94) EP <b>(34) Countries for which the regional or international application was filed:</b> AT et al. <b>(71) Applicant (for AU BB CA GB IE LK MN MW NZ SD only):</b> UNILEVER N.V. [NL/NL]; Weena 455, NL-3013 AL Rotterdam (NL). <b>(71) Applicant (for all designated States except AU BB CA GB IE LK MN MW NZ SD):</b> UNILEVER PLC [GB/GB]; Unilever House, Blackfriars, London EC4 4BQ (GB). <b>(72) Inventors:</b> FERINGA, Ben, L.; Rijksuniversiteit Groningen, Nijenborg 4, NL-9747 AG Groningen (NL). LUBBEN, Marcel; Rijksuniversiteit Groningen, Nijenborg 4, NL-9747 AG Groningen (NL). HERMANT, Roelant, Mathijs; Cyperzegge 1, NL-2318 ZR Leiden (NL). TWISKER, Robin, Stefan; Hekbootstraat 1, NL-3028 XD Rotterdam (NL). QUE, Lawrence, Jr.; 1084 Shryer Avenue West, Roseville, MN 55113 (US).		<b>(74) Common Representative:</b> UNILEVER N.V.; Patent Division, P.O. Box 137, NL-3130 AC Vlaardingen (NL). <b>(81) Designated States:</b> AM, AT, AU, BB, BG, BR, BY, CA, CH, CN, CZ, DE, DK, EE, ES, FI, GB, GE, HU, IS, JP, KE, KG, KP, KR, KZ, LK, LR, LT, LU, LV, MD, MG, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, TJ, TM, TT, UA, UG, UZ, VN, European patent (AT, BE, CH, DE, DK, ES, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG), ARIPO patent (KE, MW, SD, SZ, UG). <b>Published</b> With international search report.
<b>(54) Title:</b> BLEACH ACTIVATION <b>(57) Abstract</b> <p>A bleach and oxidation catalyst is provided comprising a catalytically active iron complex which can activate hydrogen peroxide or peroxy acids and was found to have both favourable stain removal and remarkable dye transfer inhibition properties. In addition, a considerably improved stability of these compounds in alkaline aqueous environment has been obtained, in particular at the peroxy compound concentrations generally present in the fabric washing liquor.</p>		

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## BLEACH ACTIVATION

### Field of the Invention

The invention relates to activation of bleaches employing  
5 peroxy compounds including hydrogen peroxide or hydrogen  
peroxide adducts, which liberate hydrogen peroxide in  
aqueous solution, and peroxy acids; to compounds that  
activate or catalyse peroxy compounds; to bleach  
10 compositions, including detergent bleach compositions,  
which contain a catalyst for peroxy compounds; and to  
processes for bleaching and/or washing substrates using the  
aforementioned types of compositions.

In particular, the present invention is concerned with the  
15 novel use of iron compounds as catalysts for the bleach  
activation of peroxy compounds.

### Background of the Invention.

Peroxide bleaching agents for the use in laundering have  
20 been known for many years. Such agents are effective in  
removing stains, such as tea, fruit, and wine stains, from  
clothing at or near boiling temperatures. The efficacy of  
peroxide bleaching agents drop off sharply at temperatures  
below 60°C.

25 Previous patent applications dealt with environmentally  
acceptable manganese ions and complexes. US patent  
4,728,455 discusses the use of Mn(III)-gluconate as  
peroxide bleach catalyst with high hydrolytic and oxidative  
30 stability; relatively high ratios of ligand (gluconate) to  
Mn are, however, needed to obtain the desired catalytic  
system. Moreover, the performance of these Mn-based  
catalysts is inadequate when used for bleaching in the low-  
temperature region of about 20°-40°C, and they are  
35 restricted in their efficacy to remove a wide range of  
stains.

In several patent documents, for instance EP-A-458,379, novel triazacyclononane-based manganese complexes are disclosed, which display a high catalytic oxidation activity at low temperatures, which is particularly  
5 suitable for bleaching purposes. A major improvement of the bleaching activity could be obtained by the fact that these compounds are stable under washing conditions, e.g. high alkalinity and oxidizing environment (as a result of the presence of hydrogen peroxide or peroxy acids).

10

In addition to the above-mentioned stain removal, dye transfer is a well-known problem in the art and has been addressed in various ways. For instance, an improved dye transfer inhibition has been obtained by using Fe-porphyrin  
15 and Fe-phtalocyanine complexes (see EP-A-537,381, EP-A-553,607, EP-A-538,228).

It is well known that the stability of Fe-co-ordination complexes in alkaline aqueous media in the presence of  
20 peroxide compounds is very poor; in EP-A-537,381 and EP-A-553,607, methods are disclosed for improvement in this respect.

This poor stability of Fe-co-ordination species has  
25 resulted in the necessity of very low concentrations of peroxide and, additionally, the use of polymers (see EP-A-538,228). These measures, however, only reduce the negative effects of the above-indicated poor stability to some extent and do not provide a complete solution for this  
30 problem.

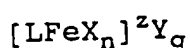
We have now surprisingly found catalytically highly active iron compounds which can activate hydrogen peroxide or peroxy acids, thereby providing both favourable stain  
35 removal, remarkable dye transfer inhibition properties, and, alternatively, oxidation of organic substrates such as olefins, alcohols and unactivated hydrocarbons.

In addition, a considerably improved stability of these compounds in alkaline aqueous environment has been obtained, in particular at the peroxy compound concentrations generally present in the wash liquor during the fabric washing cycle.

#### Definition of the Invention

In one aspect, the present invention provides a bleach and oxidation catalyst comprising an Fe-complex having formula

10 A



(A)

or precursors thereof, in which

Fe is iron in the II, III, IV or V oxidation state;

15 X represents a coordinating species such as  $\text{H}_2\text{O}$ ,  $\text{ROH}$ ,  $\text{NR}_3$ ,  $\text{RCN}$ ,  $\text{OH}^-$ ,  $\text{OOH}^-$ ,  $\text{RS}^-$ ,  $\text{RO}^-$ ,  $\text{RCOO}^-$ ,  $\text{OCN}^-$ ,  $\text{SCN}^-$ ,  $\text{N}_3^-$ ,  $\text{CN}^-$ ,  $\text{F}^-$ ,  $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{I}^-$ ,  $\text{O}^{2-}$ ,  $\text{NO}_3^-$ ,  $\text{NO}_2^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{SO}_3^{2-}$ ,  $\text{PO}_4^{3-}$  or aromatic N donors such as pyridines, pyrazines, pyrazoles, imidazoles, benzimidazoles, pyrimidines, triazoles and  
20 thiazoles with R being H, optionally substituted alkyl, optionally substituted aryl;

n is an integer number ranging from 0-3;

Y is a counter ion, the type of which is dependent on the charge of the complex;  $q = z/[\text{charge Y}]$ ;

25 z denotes the charge of the complex and is an integer which can be positive, zero or negative; if z is positive, Y is an anion such as  $\text{F}^-$ ,  $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{I}^-$ ,  $\text{NO}_3^-$ ,  $\text{BPh}_4^-$ ,  $\text{ClO}_4^-$ ,  $\text{BF}_4^-$ ,  $\text{PF}_6^-$ ,  $\text{RSO}_3^-$ ,  $\text{RSO}_4^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{CF}_3\text{SO}_3^-$ ,  $\text{RCOO}^-$  etc; if z is negative, Y is a common cation such as an alkali metal,  
30 alkaline earth metal or (alkyl)ammonium cation etc);

L represents a ligand which is an organic molecule containing a number of hetero atoms, e.g. N, P, O, S etc. which co-ordinates via all or some of its hetero atoms and/or carbon atoms to the iron centre.

35

In another aspect, the present invention provides a bleaching composition comprising a peroxy compound bleach

preferably selected from hydrogen peroxide, hydrogen peroxide- liberating or -generating compounds, peroxyacids and their salts, peroxyacid bleach precursors, and mixtures thereof, and a catalyst according to the present invention.

5

#### Detailed Description of the Invention

Generally, the Fe-complex catalyst of the invention may be used in a bleaching system comprising a peroxy compound or a precursor thereof and suitable for use in the washing and  
10 bleaching of substrates including laundry, dishwashing and hard surface cleaning. Alternatively, the Fe-complex catalyst of the invention may also be used in the textile, paper and woodpulp industries.

15 As already stated, an advantage of the Fe-complex catalysts according to the present invention is that they exhibit both a high oxidation activity and a remarkably high stability in alkaline aqueous media in the presence of peroxy compounds.

20

A second advantage of the Fe-complex catalysts of the invention is that their optimal bleaching activity is observed at lower pH values than those observed for the triazacyclononane-based manganese complex compounds  
25 mentioned above. This advantage may turn out to be very beneficial in view of the current tendency to shift the pH during fabric washing from highly alkaline (typically, a pH of 10) to more neutral values.

30 An additional advantage is that such compounds are active as dye-transfer inhibition agents as shown in Example 3. Another advantage is that the catalysts of the invention have a relatively low molecular weight and are, consequently, very weight-effective. Furthermore, they can  
35 be easily prepared.

Precursors of the active Fe-complex catalysts of the

invention can be any iron co-ordination complex, which, under fabric washing conditions, is transformed into the active iron complex of general formula A. Alternatively, the precursor of the Fe-complex of the invention can be a mixture of an iron salt, such as  $\text{Fe}(\text{NO}_3)_3$ , and the ligand L (see Example 2).

A preferred class of ligands is that of pentadentate ligands, which co-ordinate via five hetero atoms, such as nitrogen, oxygen and sulphur atoms, to the Fe atom. These hetero atoms are preferably nitrogen atoms. The nitrogen atoms can be part of tertiary, secondary or primary amine groups, tertiary, secondary or primary amide groups, or part of heterocyclic aromatic ring systems, e.g. pyridines, pyrazines, pyrazoles, imidazoles, benzimidazoles, thiazoles, triazoles and pyrimidines, or combinations thereof.

Examples of preferred ligands in their simplest forms are:

20

(i) pyridin-2-yl-methyl containing ligands such as:  
N,N-bis(pyridin-2-yl-methyl)-bis(pyridin-2-yl)methylamine;  
N,N-bis(pyrazol-1-yl-methyl)-bis(pyridin-2-yl)methylamine;  
N,N-bis(imidazol-2-yl-methyl)-bis(pyridin-2-yl)methylamine;  
25 N,N-bis(1,2,4-triazol-1-yl-methyl)-bis(pyridin-2-yl)methylamine;  
N,N-bis(pyridin-2-yl-methyl)-bis(pyrazol-1-yl)methylamine;  
N,N-bis(pyridin-2-yl-methyl)-bis(imidazol-2-yl)methylamine;  
30 N,N-bis(pyridin-2-yl-methyl)-bis(1,2,4-triazol-1-yl)methylamine;

(ii) 2-amino-ethyl containing ligands such as:  
35 N,N-bis(2-amino-ethyl)-bis(pyridin-2-yl)methylamine;  
N,N-bis(2-amino-ethyl)-bis(pyrazol-1-yl)methylamine;  
N,N-bis(2-amino-ethyl)-bis(imidazol-2-yl)methylamine;

N,N-bis(2-amino-ethyl)-bis(1,2,4-triazol-1-yl) methylamine;  
N,N-bis(pyridin-2-yl-methyl)-bis(2-amino-ethyl)methylamine;  
N,N-bis(pyrazol-1-yl-methyl)-bis(2-amino-ethyl)methylamine;  
N,N-bis(imidazol-2-yl-methyl)-bis(2-amino-

5 ethyl)methylamine;

N,N-bis(1,2,4-triazol-1-yl-methyl)-bis(2-amino-ethyl)methylamine;

The most preferred ligand is :

10 N,N-bis(pyridin-2-yl-methyl)-bis(pyridin-2-yl)methylamine,  
hereafter referred to as  $N_4Py$ .

Suitable counter ions are those which give rise to the formation of storage-stable solids. Combination of the  
15 preferred iron complexes with the counter ion Y preferably involves counter ions such as  $RCOO^-$ ,  $BPh_4^-$ ,  $ClO_4^-$ ,  $BF_4^-$ ,  $PF_6^-$ ,  $RSO_3^-$ ,  $RSO_4^-$ ,  $SO_4^{2-}$ ,  $NO_3^-$ , wherein R=H, optionally substituted phenyl, naphthyl or  $C_1-C_4$  alkyl. Preferred co-ordinating species X are  $CH_3CN$ ,  $H_2O$ ,  $Cl^-$ ,  $OH^-$ , and  $OOH^-$ .

20

The effective level of the Fe-complex catalyst, expressed in terms of parts per million (ppm) of iron in an aqueous bleaching solution, will normally range from 0.001 ppm to 100 ppm, preferably from 0.01 ppm to 20 ppm, most  
25 preferably from 0.1 ppm to 10 ppm. Higher levels may be desired and applied in industrial bleaching processes, such as textile and paper pulp bleaching. The lower range levels are preferably used in domestic laundry operations.

30

#### The detergent bleach composition

The bleaching composition of the invention has particular application in detergent formulations to form a new and improved detergent bleach composition within the purview of  
35 the invention, comprising a peroxy compound bleach as defined above, the aforesaid Fe-complex catalyst having general formula (A), a surface-active material and a



detergency builder.

The Fe-complex catalyst will be present in the detergent bleach composition of the invention in amounts so as to provide the required level in the wash liquor. Generally, the Fe-complex catalyst level in the detergent bleach composition corresponds to an iron content of from 0.0005% to 0.5% by weight. When the dosage of detergent bleach composition is relatively low, e.g. about 1-2 g/l, the Fe content in the formulation is suitably 0.0025 to 0.5%, preferably 0.005 to 0.25% by weight. At higher product dosages, as used e.g. by European consumers, the Fe-content in the formulation is suitably 0.0005 to 0.1%, preferably 0.001 to 0.05% by weight.

Detergent bleach compositions of the invention are effective over a wide pH-range of between 7 and 13, with optimal pH-range lying between 8 and 11.

#### The peroxy bleaching compound

The peroxy bleaching compound may be a compound which is capable of yielding hydrogen peroxide in aqueous solution. Hydrogen peroxide sources are well known in the art. They include the alkali metal peroxides, organic peroxides such as urea peroxide, and inorganic persalts, such as the alkali metal perborates, percarbonates, perphosphates persilicates and persulphates. Mixtures of two or more such compounds may also be suitable.

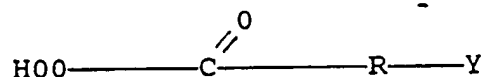
Particularly preferred are sodium perborate tetrahydrate and, especially, sodium perborate monohydrate. Sodium perborate monohydrate is preferred because of its high active oxygen content. Sodium percarbonate may also be preferred for environmental reasons. The amount thereof in the composition of the invention usually will be within the range of about 5-35 % by weight, preferably from 10-25 % by weight.

Another suitable hydrogen peroxide generating system is a combination of a C<sub>1</sub>-C<sub>4</sub> alkanol oxidase and a C<sub>1</sub>-C<sub>4</sub> alkanol, especially a combination of methanol oxidase (MOX) and ethanol (see Example 3). Such combinations are disclosed in  
5 International Application PCT/EP 94/03003 (Unilever), which is incorporated herein by reference.

Alkylhydroxy peroxides are another class of peroxy bleaching compounds. Examples of these materials include  
10 cumene hydroperoxide and t-butyl hydroperoxide.

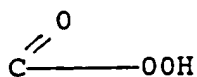
Organic peroxyacids may also be suitable as the peroxy bleaching compound. Such materials normally have the general formula:

15



wherein R is an alkylene or substituted alkylene group containing from 1 to about 20 carbon atoms, optionally  
20 having an internal amide linkage; or a phenylene or substituted phenylene group; and Y is hydrogen, halogen, alkyl, aryl, an imido-aromatic or non-aromatic group, a COOH or

25



group or a quaternary ammonium group.

Typical monoperoxy acids useful herein include, for  
30 example:

- (i) peroxybenzoic acid and ring-substituted peroxybenzoic acids, e.g. peroxy- $\alpha$ -naphthoic acid;
- (ii) aliphatic, substituted aliphatic and arylalkyl  
35 monoperoxyacids, e.g. peroxy lauric acid, peroxy stearic acid and N,N-phthaloylaminoperoxy caproic acid (PAP); and

(iii) 6-octylamino-6-oxo-peroxyhexanoic acid.

Typical diperoxyacids useful herein include, for example:

- 5 (iv) 1,12-diperoxydodecanedioic acid (DPDA);
- (v) 1,9-diperoxyazelaic acid;
- (vi) diperoxybrassicilic acid; diperoxysebasic acid and  
diperoxyisophthalic acid;
- (vii) 2-decyldiperoxybutane-1,4-dioic acid; and
- 10 (viii) 4,4'-sulphonylbisperoxybenzoic acid.

Also inorganic peroxyacid compounds are suitable, such as  
for example potassium monopersulphate (MPS). If organic or  
inorganic peroxyacids are used as the peroxygen compound,  
15 the amount thereof will normally be within the range of  
about 2-10 % by weight, preferably from 4-8 % by weight.

All these peroxy compounds may be utilized alone or in  
conjunction with a peroxyacid bleach precursor and/or an  
20 organic bleach catalyst not containing a transition metal.  
Generally, the bleaching composition of the invention can  
be suitably formulated to contain from 2 to 35% ,  
preferably from 5 to 25% by weight, of the peroxy bleaching  
agent.

25

Peroxyacid bleach precursors are known and amply described  
in literature, such as in the British Patents 836988;  
864,798; 907,356; 1,003,310 and 1,519,351; German Patent  
3,337,921; EP-A-0185522; EP-A-0174132; EP-A-0120591; and US  
30 Patents 1,246,339; 3,332,882; 4,128,494; 4,412,934 and  
4,675,393.

Another useful class of peroxyacid bleach precursors is  
that of the cationic i.e. quaternary ammonium substituted  
35 peroxyacid precursors as disclosed in US Patent 4,751,015  
and 4,397,757, in EP-A0284292 and EP-A-331,229. Examples of  
peroxyacid bleach precursors of this class are:

2-(N,N,N-trimethyl ammonium) ethyl sodium-4-sulphonphenyl  
carbonate chloride - (SPCC);

N-octyl,N,N-dimehyl-N<sub>10</sub>-carbophenoxy decyl ammonium  
chloride - (ODC);

- 5 3-(N,N,N-trimethyl ammonium) propyl sodium-4-sulphophenyl  
carboxylate; and  
N,N,N-trimethyl ammonium toluyloxy benzene sulphonate.

A further special class of bleach precursors is formed by  
10 the cationic nitriles as disclosed in EP-A-303,520 and in  
European Patent Specification No.'s 458,396 and 464,880.

Any one of these peroxyacid bleach precursors can be used  
in the present invention, though some may be more preferred  
15 than others.

Of the above classes of bleach precursors, the preferred  
classes are the esters, including acyl phenol sulphonates  
and acyl alkyl phenol sulphonates; the acyl-amides; and the  
20 quaternary ammonium substituted peroxyacid precursors  
including the cationic nitriles.

Examples of said preferred peroxyacid bleach precursors or  
activators are sodium-4-benzoyloxy benzene sulphonate  
25 (SBOBS); N,N,N,N'-tetraacetyl ethylene diamine (TAED);  
sodium-1-methyl-2-benzoyloxy benzene-4-sulphonate; sodium-  
4-methyl-3-benzoloxo benzoate; SPCC; trimethyl ammonium  
toluyloxy-benzene sulphonate; sodium nonanoyloxybenzene  
sulphonate (SNOBS); sodium 3,5,5-trimethyl hexanoyl-  
30 oxybenzene sulphonate (STHOBS); and the substituted  
cationic nitriles.

The precursors may be used in an amount of up to 12 %, preferably from 2-10 % by weight, of the composition.

35

As an alternative to the above described peroxide  
generating systems, molecular oxygen may be used as the

oxidant.

The surface-active material

The detergent bleach composition according to the present invention generally contains a surface-active material in an amount of from 10 to 50% by weight. Said surface-active material may be naturally derived, such as soap, or a synthetic material selected from anionic, nonionic, amphoteric, zwitterionic, cationic actives and mixtures thereof. Many suitable actives are commercially available and are fully described in the literature, for example in "Surface Active Agents and Detergents", Volumes I and II, by Schwartz, Perry and Berch.

Typical synthetic anionic surface-actives are usually water-soluble alkali metal salts of organic sulphates and sulphonates having alkyl radicals containing from about 8 to about 22 carbon atoms, the term alkyl being used to include the alkyl portion of higher aryl radicals. Examples of suitable synthetic anionic detergent compounds are sodium and ammonium alkyl sulphates, especially those obtained by sulphating higher (C<sub>8</sub>-C<sub>18</sub>) alcohols produced, for example, from tallow or coconut oil; sodium and ammonium alkyl (C<sub>9</sub>-C<sub>10</sub>) benzene sulphonates, particularly sodium linear secondary alkyl (C<sub>10</sub>-C<sub>15</sub>) benzene sulphonates; sodium alkyl glyceryl ether sulphates, especially those ester of the higher alcohols derived from tallow or coconut oil fatty acid monoglyceride sulphates and sulphonates; sodium and ammonium salts of sulphuric acid esters of higher (C<sub>9</sub>-C<sub>18</sub>) fatty alcohol alkylene oxide, particularly ethylene oxide, reaction products; the reaction products of fatty acids such as coconut fatty acids esterified with isethionic acid and neutralised with sodium hydroxide; sodium and ammonium salts of fatty acid amides of methyl taurine; alkane monosulphonates such as those derived by reacting alpha-olefins (C<sub>8</sub>-C<sub>20</sub>) with sodium bisulphite and those derived by reaction paraffins with SO<sub>2</sub>

- and C<sub>12</sub> and then hydrolysing with a base to produce a random sulphonate; sodium an ammonium C<sub>7</sub>-C<sub>12</sub> dialkyl sulphosccinates; and olefin sulphonates which term is used to describe material made by reacting olefins, particularly
- 5 C<sub>10</sub>-C<sub>20</sub> alpha-olefins, with SO<sub>3</sub> and then neutralising and hydroysing the reaction product. The preferred anionic detergent compounds are sodium (C<sub>10</sub>-C<sub>15</sub>) alkylbenzene sulphonates, sodium C<sub>16</sub>-C<sub>18</sub>) alkyl ether sulphates.
- 10 Examples of suitable nonionic surface-active compounds which may be used, preferably together with the anionic surface-active compounds, include, in particular, the reaction products of alkylene oxides, usually ethylene oxide, with alkyl (C<sub>6</sub>-C<sub>22</sub>) phenols, generally 5-25 EO, i.e.
- 15 5-25 units of ethylene oxides per molecule; and the condensation products of aliphatic (C<sub>8</sub>-C<sub>18</sub>) primary or secondary linear or branched alcohols with ethylene oxide, generally 2-30 EO. Other so-called nonionic surface-
- 20 chain tertiary amine oxides, long-chain tertiary phosphine oxides and dialkyl sulfoxides.

Amphoteric or zwitterionic surface-active compounds can also be used in the compositions of the invention but this

25 is not normally desired owing to their relatively high cost. If any amphoteric or zwitterionic detergent compounds are used, it is generally in small amounts in compositions based on the much more commonly used synthetic anionic and nonionic actives.

30

As disclosed by EP-A-544,490, the performance of the hereinbefore described bleach catalyst, may be dependent upon the active detergent system and the builder system present in the detergent bleach composition of the

35 invention.

The detergent bleach composition of the invention will

preferably comprise from 1-15 % wt of anionic surfactant and from 10-40 % by weight of nonionic surfactant. In a further preferred embodiment the detergent active system is free from C<sub>16</sub>-C<sub>12</sub> fatty acids soaps.

5

The detergency builder

The composition of the invention normally and preferably also contains a detergency builder in an amount of from about 5-80 % by weight, preferably from about 10-60 % by weight.

10

Builder materials may be selected from 1) calcium sequestrant materials, 2) precipitating materials, 3) calcium ion-exchange materials and 4) mixtures thereof.

15

Examples of calcium sequestrant builder materials include alkali metal polyphosphates, such as sodium tripolyphosphate; nitrilotriacetic acid and its water-soluble salts; the alkali metal salts of carboxymethyloxy succinic acid, ethylene diamine tetraacetic acid, oxydisuccinic acid, mellitic acid, benzene polycarboxylic acids, citric acid; and polyacetal carboxylates as disclosed in US Patents 4,144,226 and 4,146,495.

20

Examples of precipitating builder materials include sodium orthophosphate and sodium carbonate.

25

Examples of calcium ion-exchange builder materials include the various types of water-insoluble crystalline or amorphous aluminosilicates, of which zeolites are the best known representatives, e.g. zeolite A, zeolite B (also known as Zeolite P), zeolite C, zeolite X, zeolite Y and also the zeolite P type as described in EP-A-0384070.

30

In particular, the compositions of the invention may contain any one of the organic and inorganic builder materials, though, for environmental reasons, phosphate

35

builders are preferably omitted or only used in very small amounts.

Typical builders usable in the present invention are , for example, sodium carbonate, calcite/carbonate, the sodium  
5 salt of nitrilotriacetic acid, sodium citrate, carboxymethyloxy malonate, carboxymethyloxy succinate and the water-insoluble crystalline or amorphous aluminosilicate builder material, each of which can be used as the main builder, either alone or in admixture with  
10 minor amounts of other builders or polymers as co-builder.

It is preferred that the composition contains not more than 5% by weight of a carbonate builder, expressed as sodium carbonate, more preferable not more than 2.5 % by weight to  
15 substantially nil, if the composition pH lies in the lower alkaline region of up to 10.

#### Other ingredients

Apart from the components already mentioned, the detergent  
20 bleach composition of the invention can contain any of the conventional additives in amounts of which such materials are normally employed in fabric washing detergent compositions. Examples of these additives include buffers such as carbonates, lather boosters, such as alkanolamides,  
25 particularly the monoethanol amides derived from palmkernel fatty acids and coconut fatty acids; lather depressants, such as alkyl phosphates and silicones; anti-redeposition agents, such as sodium carboxymethyl cellulose and alkyl or substituted alkyl cellulose ethers; stabilizers, such as  
30 phosphonic acid derivatives (i.e. Dequest® types); fabric softening agents; inorganic salts and alkaline buffering agents, such as sodium sulphate, sodium silicate etc.; and usually in very small amounts, fluorescent agents; perfumes; enzymes, such as proteases, cellulases, lipases,  
35 amylases and oxidases; germicides and colourants.

When using a hydrogenperoxide source, such as sodium



perborate or sodium percarbonate, as the bleaching compound, it is preferred that the composition contains not more than 5 % by weight of a carbonate buffer, expressed as sodium carbonate, more preferable not more than 2.5% by weight to substantially nil, if the composition pH lies in the lower alkaline region of up to 10.

Of the additives, transition metal sequestrants, such as EDTA and the phosphonic acid derivatives, e.g. ethylene diamine tetra-(methylene phosphonate)-EDTMP- are of special importance, as not only do they improve the stability of the catalyst/H<sub>2</sub>O<sub>2</sub> system and sensitive ingredients, such as enzymes, fluorescent agents, perfumes and the like, but also improve the bleach performance, especially at the higher pH region of above 10, particularly at pH 10.5 and above.

The invention will now be further illustrated by way of the following non-limiting Examples.

20

Example 1

Preparation of [Fe(N<sub>4</sub>Py)(CH<sub>3</sub>CN)](ClO<sub>4</sub>)<sub>2</sub>·2 H<sub>2</sub>O (FeL<sub>1</sub>).

The ligand N<sub>4</sub>Py was prepared as follows:

25 To pyridyl ketone oxim (3 g, 15.1 mmol) was added ethanol (15 ml), concentrated ammonia solution (15 mL) and NH<sub>4</sub>OAc (1.21 g, 15.8 mmol). The solution was warmed until reflux. To this solution was added 4.64 g Zn in small portions. After the addition of all Zn, the mixture was refluxed for 30 1 h. and allowed to cool to ambient temperature. The solution was filtered and water (15 ml) was added. Solid NaOH was added until PH>>10 and the solution was extracted with CH<sub>2</sub>Cl<sub>2</sub> (3 x 20 ml). The organic layers were dried over Na<sub>2</sub>SO<sub>4</sub> and evaporated until dryness. Bis(pyridin-2-yl)methylamine (2.39 g, 12.9 mmol) was obtained as a 35 colourless oil in 86% yield, showing the following analytical characteristics:

$^1\text{H}$  NMR (360 MHz,  $\text{CDCl}_3$ ):  $\delta$  2.64 (s, 2H,  $\text{NH}_2$ ), 5.18 (s, 1H, CH), 6.93 (m, 2H, pyridine), 7.22 (m, 2H, pyridine), 7.41 (m, 2H, pyridine), 8.32 (m, 2H, pyridine);  
 $^{13}\text{C}$  NMR ( $\text{CDCl}_3$ ):  $\delta$  62.19 (CH), 121.73 (CH), 122.01 (CH),  
5 136.56 (CH), 149.03 (CH), 162.64 (Cq).

To picolylchloride hydrochloride (4.06 g, 24.8 mmol) was added, at  $0^\circ\text{C}$ , 4.9 ml of a 5N NaOH solution. This emulsion was added by means of a syringe to bis(pyridin-2-yl)methylamine (2.3 g, 12.4 mmol) at  $0^\circ\text{C}$ . Another 5 ml of a  
10 5N NaOH solution was added to this mixture. After warming to ambient temperature, the mixture was stirred vigorously for 40 hrs. The mixture was put in an ice bath and  $\text{HClO}_4$  was added until  $\text{pH} < 1$ , whereupon a brown solid precipitated.  
15 The brown precipitate was collected by filtration and recrystallized from water. While stirring, this mixture was allowed to cool to ambient temperature, whereupon a light-brown solid precipitated which was collected by filtration and washed with cold water and air-dried (1.47 g). The free  
20 amine could be obtained by precipitating the salt with 2N NaOH and extraction with  $\text{CH}_2\text{Cl}_2$ . The free amine should be stored under an atmosphere of argon because it is sensitive to  $\text{CO}_2$ .

25 The ligand  $\text{N}_4\text{Py}$  showed the following characteristics:  
 $^1\text{H}$  NMR (360 MHz,  $\text{CDCl}_3$ ):  $\delta$  3.96 (s, 4H,  $\text{CH}_2$ ), 5.34 (s, 1H, CH), 7.00-7.10 (m, 4H, pyridine), 7.52-7.64 (m, 8H, pyridine), 8.44-8.53 (m, 4H, pyridine);  
 $^{13}\text{C}$  NMR ( $\text{CDCl}_3$ ):  $\delta$  57.36 ( $\text{CH}_2$ ), 72.06 (CH), 121.90 (CH),  
30 122.18 (CH), 123.01 (CH), 124.09 (CH), 136.35 (CH), 136.46 (CH), 149.06 (CH), 149.33 (CH), 159.83 (Cq), 160.00 (Cq).

Subsequently,  $[(\text{N}_4\text{Py})\text{Fe}(\text{CH}_3\text{CN})](\text{ClO}_4)_2 \cdot 2 \text{H}_2\text{O}$  was prepared as follows:

35

To a solution of 0.144 g (0.392 mmol) of  $\text{N}_4\text{Py}$  in methanol/acetonitrile (2ml/2ml) was added 0.215 g (0.403

- mmol)  $\text{Fe}(\text{ClO}_4)_3 \cdot 10 \text{H}_2\text{O}$ . After stirring for 5 minutes, the solution was placed in an ethyl acetate bath. The crystals formed overnight, were collected and washed with ethyl acetate to yield 0.178 grams of  $[(\text{N}_4\text{Py})\text{Fe}(\text{CH}_3\text{CN})](\text{ClO}_4)_2 \cdot 2 \text{H}_2\text{O}$  as dark red crystals (yield: 65%).
- 5  $^1\text{H}$  NMR (360 MHz,  $\text{CD}_3\text{CN}$ ):  $\delta$  4.27 (d, 2H,  $J=18.1$  Hz), 4.40 (d, 2H,  $J=18.1$  Hz), 6.34 (s, 1H); 7.06 (d, 2H,  $J=7.8$  Hz), 7.33 (m, 4H), 7.68 (m, 2H), 7.88 (m, 4H), 8.90 (d, 2H,  $J=5.4$  Hz), 9.03 (d, 2H,  $J=5.4$  Hz)
- 10 UV-vis (acetone) [ $\lambda_{\text{max}}$ , nm ( $\epsilon$ ,  $\text{M}^{-1}\text{cm}^{-1}$ ): 382 (5650), 458 (3970)
- Anal. Calcd for  $\text{C}_{25}\text{H}_{28}\text{Cl}_2\text{FeN}_6\text{O}_{10}$ : C, 42.94; H, 4.04; N, 12.02. Found: C, 43.21; H, 3.76; N, 12.02.

- 15 In the Examples 2, 3 and 4, the above-described complex  $[(\text{N}_4\text{Py})\text{Fe}(\text{CH}_3\text{CN})](\text{ClO}_4)_2 \cdot 2 \text{H}_2\text{O}$  is referred to as  $\text{Fe}(\text{N}_4\text{Py})$ .

#### Example 2

- The bleaching activity of the Fe-catalyst prepared
- 20 according to Example 1 was demonstrated in the presence hydrogen peroxide on standard tea-stained (BC-1) cotton test cloths.

- The experiments were carried out at  $40^\circ\text{C}$  and at a pH of 6
- 25 and 8 in a temperature-controlled glass beaker equipped with a magnetic stirrer, thermocouple and a pH electrode.

- Two pieces of test cloth were stirred for 60 minutes in 1 liter of a  $8.6 \times 10^{-3}$  mol/l hydrogen peroxide solution in
- 30 millipore water, containing concentrations of the compounds as indicated in Table 1. After rinsing with demineralised water, the test cloths were dried for 7 minutes in a microwave oven. The reflectance ( $R_{460^*}$ ) of the test cloths was measured on a Macbeth 1500/plus colour measuring system
- 35 including UV-filter before and after treatment. The difference ( $\Delta R_{460^*}$ ) between both reflectance values thus obtained gives a measure of the bleaching performance, i.e.

higher  $\Delta R_{460}^*$  values correspond to an improved bleaching performance.

TABLE 1

5	conc.	$\Delta R_{460}^*$	$\Delta R_{460}^*$
	(mol/l)	(at pH=6)	(at pH=8)
	-	4.4	4.4
	$\text{Fe}(\text{NO}_3)_3$	3.2	4.6
	$\text{Fe}(\text{N}_4\text{Py})$	7.9	12.4
10	$\text{Fe}(\text{NO}_3)_3 + \text{N}_4\text{Py}$ $5 \times 10^{-6} (\text{Fe}) +$ $10 \times 10^{-6} (\text{N}_4\text{Py})$	-	7.5

In Table 1,  $\text{Fe}(\text{N}_4\text{Py})$  refers to the Fe-catalyst prepared according to Example 1. The blank and  $\text{Fe}(\text{NO}_3)_3$  experiment  
15 were used as control. As observed in Table 1, no experiment has been carried out at pH=6 for  $\text{Fe}(\text{NO}_3)_3 + \text{N}_4\text{Py}$ .

These measurements show that improved bleaching performance is obtained when  $\text{Fe}(\text{N}_4\text{Py})$  or a combination of Fe and the  
20 ligand  $\text{N}_4\text{Py}$  are present in solution.

### Example 3

The bleaching activity of the Fe catalyst, prepared according to example 1, was demonstrated in the presence of  
25 the hydrogen peroxide generating enzyme Methanol Oxidase (MOX) and ethanol on standard tea stained (BC-1) cotton test cloths.

The experiments were carried out in a phosphate buffer  
30 (pH=8) at 40°C in a temperature controlled glass beaker equipped with a magnetic stirrer, thermocouple and a pH electrode.

Two pieces of test cloth were vigorously shaken for 6 hours  
35 in 0.25 liter of an oxygen saturated phosphate buffer at pH=8 (millipore water), containing 200 mg MOX, 20 mM ethanol and concentrations of the compounds as indicated in

Table 2. As described in example 2, higher  $\Delta R_{460}^*$  values correspond to an improved bleaching performance.

TABLE 2

	conc. (mol/l)	$\Delta R_{460}^*$
blanc	-	7.2
$\text{Fe}(\text{NO}_3)_3$	$1 \times 10^{-5}$	6.8
$\text{Fe}(\text{N}_4\text{Py})$	$1 \times 10^{-5}$	10.8

10

In Table 2,  $\text{Fe}(\text{N}_4\text{Py})$  refers to the Fe catalyst, prepared according to example 1. The blanc and  $\text{Fe}(\text{NO}_3)_3$  experiment were used as control.

15 These measurements show that improved bleaching performance is obtained when  $\text{Fe}(\text{N}_4\text{Py})$  is present in solution.

Example 4

20 The dye oxidation activity of the Fe-catalyst prepared according to Example 1 was demonstrated in the presence peracetic acid on a dye known as acid red 88.

The experiments were carried out at ambient temperature at pH=8 in a 1 cm cuvet in the presence of  $2.2 \times 10^{-3}$  mol/l peracetic acid and  $2 \times 10^{-4}$  mol/l acid red 88. The absorbance at 503 nm ( $A_{503}$ ), which is the maximum of the characteristic visible absorption of the dye in aqueous media, was measured at t=0 and t=15 minutes. The ratio ( $\Delta A_{503} = A_{503}(t=15) / A_{503}(t=0 \text{ min})$ ) of the absorbance at t=15 minutes and t=0 gives a measure of the dye-oxidation performance, i.e. an improved dye-oxidation performance results in reduced  $\Delta A_{503}$  values.

25

30

TABLE 3

	conc. (mol/l)	$\Delta A_{503}$
blanc	-	1.00
5 Fe(NO <sub>3</sub> ) <sub>3</sub>	$5 \times 10^{-6}$	0.98
Fe(N <sub>4</sub> Py)	$5 \times 10^{-6}$	0.075

Fe(N<sub>4</sub>Py) in Table 2 refers to the Fe-catalyst prepared according to Example 1. The blanc and Fe(NO<sub>3</sub>)<sub>3</sub> experiment  
10 were used as controls.

These measurements show that improved dye oxidation performance is obtained when Fe(N<sub>4</sub>Py) is present in solution.

15

Example 5

The organic substrate oxidation activity of the Fe catalyst, prepared according to example 1, was demonstrated in the presence of hydrogen peroxide on a range of organic  
20 substrates.

The experiments were carried out at ambient temperature in acetone. The concentration of the Fe catalyst was  $3.5 \times 10^{-6}$  M and the ratio catalyst/H<sub>2</sub>O<sub>2</sub>/substrate was 1/100/860. The turnover numbers indicated in Table 4 represent the number  
25 of molecules formed per molecule of the catalyst as determined after the indicated time of reaction by using gas chromatography. In a blanc experiment or in the presence of Fe(NO<sub>3</sub>)<sub>3</sub>, essentially no oxidation products could be detected.

30

TABLE 4

	substrate	product (turnover number)	reaction time
5			
	cyclohexene	2-cyclohexen-1-ol (18) 2-cyclohexen-1-one (9) cyclohexene epoxide (2)	30 minutes
10	cyclohexane	cyclohexanol (13) cyclohexanon (5)	30 minutes
	benzylalcohol	benzylaldehyde (38)	30 minutes
	styrene	benzylaldehyde (23)	30 minutes
15			
	adamantane	1-adamantanol (7) 2-adamantanol (7) 2-adamantanone (4)	60 minutes

CLAIMS

1. A bleach and oxidation catalyst comprising an Fe-complex having formula A

5



or precursors thereof, in which

Fe is iron in the II, III, IV or V oxidation state;

10 X represents a co-ordinating species such as  $\text{H}_2\text{O}$ ,  $\text{ROH}$ ,  $\text{NR}_3$ ,  $\text{RCN}$ ,  $\text{OH}^-$ ,  $\text{OOH}^-$ ,  $\text{RS}^-$ ,  $\text{RO}^-$ ,  $\text{RCOO}^-$ ,  $\text{OCN}^-$ ,  $\text{SCN}^-$ ,  $\text{N}_3^-$ ,  $\text{CN}^-$ ,  $\text{F}^-$ ,  $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{I}^-$ ,  $\text{O}^{2-}$ ,  $\text{NO}_3^-$ ,  $\text{NO}_2^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{SO}_3^{2-}$ ,  $\text{PO}_4^{3-}$  or aromatic N donors such as pyridines, pyrazines, pyrazoles, imidazoles, benzimidazoles, pyrimidines, triazoles and  
15 thiazoles with R being H, optionally substituted alkyl, optionally substituted aryl;

n is an integer number ranging from 0-3;

Y is a counter ion, the type of which is dependent on the charge of the complex;

20 z denotes the charge of the complex and is an integer which can be positive, zero or negative; if z is positive, Y is an anion such as  $\text{F}^-$ ,  $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{I}^-$ ,  $\text{NO}_3^-$ ,  $\text{BPh}_4^-$ ,  $\text{ClO}_4^-$ ,  $\text{BF}_4^-$ ,  $\text{PF}_6^-$ ,  $\text{RSO}_3^-$ ,  $\text{RSO}_4^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{CF}_3\text{SO}_3^-$ ,  $\text{RCOO}^-$  etc; if z is negative, Y is a common cation such as an alkali metal,  
25 alkaline earth metal or (alkyl)ammonium cation etc);

$q = z / [\text{charge Y}]$ ;

L represents a ligand which is an organic molecule containing a number of hetero atoms, e.g. N, P, O, S etc. which co-ordinates via all or some of its hetero atoms  
30 and/or carbon atoms to the iron centre.

2. Catalyst according Claim 1, wherein L represents a pentadentate ligand which co-ordinates via five nitrogen atoms to the Fe atom.



3. Catalyst according to Claim 1 or 2, wherein X represents a co-ordinating species selected from  $\text{CH}_3\text{CN}$ ,  $\text{H}_2\text{O}$ ,  $\text{Cl}^-$ ,  $\text{OH}^-$ , and  $\text{OOH}^-$ .
- 5 4. Catalyst according to any of Claims 1-3, wherein the counter ion Y is selected from  $\text{RCOO}^-$ ,  $\text{BPh}_4^-$ ,  $\text{ClO}_4^-$ ,  $\text{BF}_4^-$ ,  $\text{PF}_6^-$ ,  $\text{RSO}_3^-$ ,  $\text{RSO}_4^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{NO}_3^-$ , wherein  $\text{R}=\text{H}$ , optionally substituted phenyl, naphthyl or  $\text{C}_1\text{-C}_4$  alkyl.
- 10 5. Catalyst according to any of Claims 1-4, wherein the ligand L is N,N-bis(pyridin-2-yl-methyl)-bis(pyridin-2-yl)methylamine.
6. A bleaching composition comprising a peroxy bleaching  
15 compound and a catalyst according to any of the preceding claims 1-5.
7. Composition according to claim 6, which comprises said  
20 peroxy bleaching compound at a level of from 2 to 35% by weight and said catalyst at a level corresponding to an iron content of from 0.0005 to 0.5% by weight.
8. Composition according to claim 6 or 7, wherein the  
25 peroxy bleaching compound is selected from the group consisting of hydrogen peroxide, hydrogen peroxide-liberating or -generating compounds, peroxyacids and their salts, peroxyacid bleach precursors, and mixtures thereof.
9. Composition according to any of claims 6-8, which  
30 further comprises a surface-active material, in an amount of from 10 to 50% by weight, and a detergency builder in an amount of from 5 to 80% by weight.

## INTERNATIONAL SEARCH REPORT

Internat. Application No

PCT/EP 95/01930

## A. CLASSIFICATION OF SUBJECT MATTER

IPC 6 C11D3/39 C07F15/02 B01J31/18 D06L3/02 D21C9/16

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 C11D C07F B01J D06L D21C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	JOURNAL OF COORDINATION CHEMISTRY, vol.14, no.3, 1986, GB pages 191 - 200 L. MAZZARELLA 'Secondary Enamines as Ligands' see page 196 see page 192 - page 193 ---	1,4
X	US,A,4 169 092 (J.BAYER) 25 September 1979 see claims 1,18; example 3 ---	1
X	EP,A,0 237 111 (UNILEVER) 16 September 1987 cited in the application see page 14 - page 15 ---	1,4,6-9
	--- -/--	



Further documents are listed in the continuation of box C.



Patent family members are listed in annex.

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PCT/EP 95/01930

## C(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	FR,A,2 073 596 (HALCON INTERNATIONAL) 1 October 1971 see examples ---	1,3
X	GB,A,1 182 143 (BORAX AND CHEMICAL) 25 February 1970 see page 2; examples 20-21 ---	1,4,6-9
X	FR,A,2 335 593 (UNILEVER) 15 July 1977 see page 3; claim 1; tables I,III,VI ---	1,3,6-9
X	EP,A,0 392 592 (UNILEVER) 17 October 1990 see page 6; claims; example VI ---	1,4,6-9
X	TEXTILE RESEARCH JOURNAL, vol.58, no.4, April 1988, PRINCETON, US pages 198 - 210 J. RUCKER, D. CATES '2,2'-Bipyridine catalysed bleaching of cotton fibers with peracetic acid' see page 207 - page 208 ---	1,4,6-8
X	US,A,5 288 746 (K.PRAMOD) 22 February 1994 see column 7; claims ---	1,6-9
A	WO,A,93 14093 (BOARD OF REGENDS, THE UNIVERSITY OF TEXAS) 22 July 1993 see page 14; claims 1,8 -----	1,2

## INTERNATIONAL SEARCH REPORT

Information on patent family members

Internat. Application No

PCT/EP 95/01930

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
US-A-4169092	25-09-79	NONE	
EP-A-0237111	16-09-87	US-A- 4728455	01-03-88
		AU-B- 595586	05-04-90
		AU-A- 6963787	10-09-87
		CA-A- 1281024	05-03-91
		DE-A- 3783626	04-03-93
		ES-T- 2053519	01-08-94
		JP-B- 6033431	02-05-94
		JP-A- 62225598	03-10-87
		JP-B- 6070240	07-09-94
		JP-A- 6088099	29-03-94
FR-A-2073596	01-10-71	BE-A- 760082	09-06-71
		CH-A- 521385	15-04-72
		DE-A- 2061218	16-06-71
		GB-A- 1336586	07-11-73
		LU-A- 62206	23-08-72
		NL-A- 7017292	15-06-71
		US-A- 3644455	22-02-72
GB-A-1182143	25-02-70	BE-A- 694721	28-08-67
		CH-A- 497595	15-10-70
		DE-A- 1594877	06-05-71
		FR-A- 1515338	17-05-68
		NL-A- 6703189	04-09-67
		US-A- 3532634	06-10-70
FR-A-2335593	15-07-77	GB-A- 1565807	23-04-80
		AT-B- 362034	27-04-81
		BE-A- 849540	17-06-77
		CA-A- 1079007	10-06-80
		CH-A- 624709	14-08-81
		DE-A- 2657043	23-06-77
		JP-C- 1029487	22-01-81
		JP-A- 52076281	27-06-77
		JP-B- 55020520	03-06-80
		NL-A- 7614001	21-06-77
		SE-B- 425319	20-09-82
		SE-A- 7614253	19-06-77

# INTERNATIONAL SEARCH REPORT

Information on patent family members

Internat. Application No

PCT/EP 95/01930

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
FR-A-2335593		US-A- 4119557	10-10-78
EP-A-0392592	17-10-90	CA-A, C 2014321	13-10-90
		DE-D- 69014384	12-01-95
		DE-T- 69014384	20-04-95
		ES-T- 2066099	01-03-95
		US-A- 5114611	19-05-92
US-A-5288746	22-02-94	EP-A- 0603931	29-06-94
		JP-A- 6234997	23-08-94
WO-A-9314093	22-07-93	US-A- 5252720	12-10-93
		AU-B- 3436793	03-08-93
		CA-A- 2127530	22-07-93
		EP-A- 0623134	09-11-94
		FI-A- 943445	20-07-94
		JP-T- 7503009	30-03-95
		NO-A- 942719	19-09-94
		US-A- 5432171	11-07-95
		US-A- 5439570	08-08-95

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